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ALPHA-METHYLSTYRL- AND STYRLPHOSPHAZENE MONOMERS AND  
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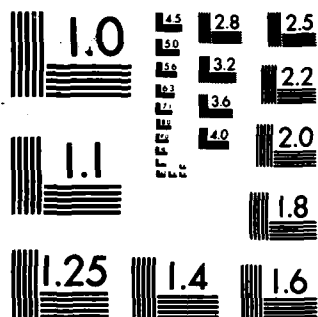
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$\alpha$ -Methylstyryl- and Styrylphosphazene Monomers and Polymers

by

Christopher W. Allen and Jonathan C. Shaw

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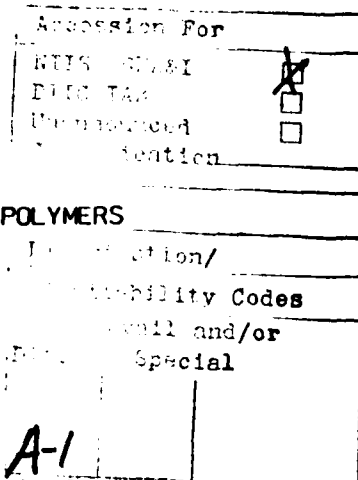
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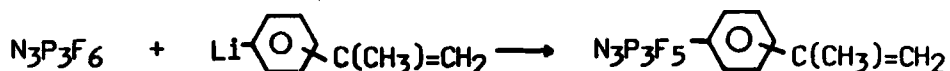
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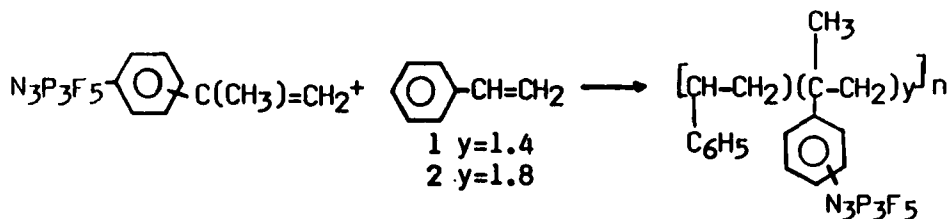
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The first system to be investigated was the  $\alpha$ -methylstyrene system. Both the meta (1) and para (2)  $\alpha$ -methylstyrylpentafluorocyclotriphosphazenes were prepared from hexafluorocyclotriphosphazene and the appropriate lithium reagent. All new materials were characterized by elemental analysis, mass spectrometry, ir and nmr ( $^1\text{H}$ ,  $^{13}\text{C}$ ,  $^{19}\text{F}$ ,  $^{31}\text{P}$ ) spectroscopy. The chemical



shifts of the vinylidene,  $=\text{CH}_2$  protons are identical in 1 and 2 suggesting the absence of significant mesomeric electron withdrawal from the aryl unit by the phosphazene. The difference in the  $\beta$ -vinyl carbon chemical shifts fall between those observed in systems with significant mesomeric interactions and those where such interactions cannot occur. This observation indicates either low degrees of mesomeric interaction or the more likely effect of electrostatic stabilization of aryl charge density by the positive phosphorus center. At the stage of bis substitution, all three isomers of  $\text{N}_3\text{P}_3\text{F}_4[\text{C}_6\text{H}_4\text{C}(\text{CH}_3)=\text{CH}_2]_2$  are formed. These systems were characterized by GC-mass spectrometry and nmr spectroscopy. The ratio of geminal to cis-non-geminal to trans-non-geminal isomers is 1:6.1:1.5 for the para- $\alpha$ -methylstyryl derivatives and 1:2.8:1.1 in the meta system. Although the predominance of non-geminal isomers can be ascribed to steric effects, the preferential formation of the cis isomer is counter-intuitive. This observation may be rationalized by postulating an electrostatic attraction between the electron poor aryl ring on the phosphazene and the electron rich aryl ring in the lithium reagent thus favoring a cis attack. The decreased amount of the cis isomer in the meta derivative can be ascribed to the steric effect of the meta substituent.

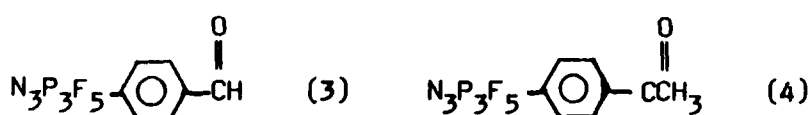
Both 1 and 2 enter into facile copolymerization reactions with styrene or para-chlorostyrene. The phosphazene incorporation in the copolymer is the highest observed for alkenyl-



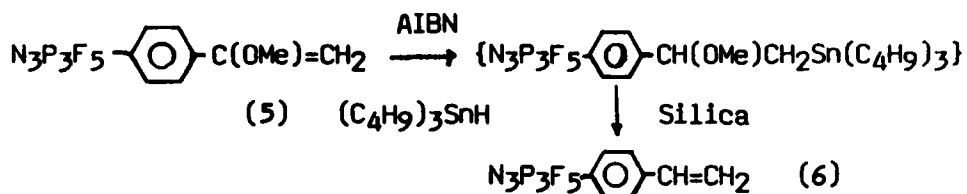
phosphazene monomers. The maximum degrees of incorporation achieved requires that there are regions of at least two phosphazene units linked together. This structural feature has not been previously observed in organophosphazene copolymers. Both of these observations are related to the moderation of the phosphazene induced olefin polarity by the aryl spacer group interposed between the two functionalities in question. Semiquantitative information on the magnitude of this effect has been obtained from the reactivity ratios and Alfrey-Price (Q,e) parameters for 1 and 2. The reactivity ratios ( $r_2$ ) of 1 and 2 as calculated by the Mortimer-Tidwell non-linear least squares approach are within the 95% joint confidence limits of each other suggesting similar

reactivity. The reactivity ratio for styrene ( $r_1$ ) as well as  $r_2$  are both less than one indicating an alternative tendency. This suggests the possibility of an electrostatic interaction between the nucleophilic styryl group and the electrophilic phosphazene. The calculated Q values for 1 and 2 are reduced from styrene showing a reduction in mesomeric interactions. The calculated e values are significantly higher than that of styrene or nitrostyrene reflecting the highly polar nature of 1 and 2. The thermal decomposition of the copolymers of 1 and 2 when studied by TGA show a two step unzipping process corresponding to phosphazene and styrene loss respectively as the temperature increases. The decomposition of the copolymers of 1 begins before similar event for copolymers of 2 indicating a role for steric destabilization in the meta substituted derivatives. The copolymers of 1 and 2 with methyl methacrylate were also obtained. Attempts to calculate reactivity ratios lead to solutions with very large 95% joint confidence limits in the phosphazene monomer. We were, however, able to achieve a good fit of the data using a penultimate model. Thus, in these systems, the reactivity of the terminal radical is significantly modified by the presence of a phosphazene function in the penultimate position.

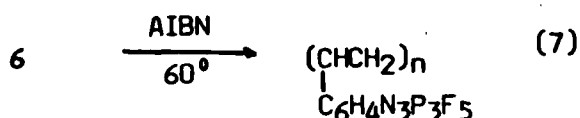
Although a broad spectrum of copolymers could be prepared from 1 or 2, all attempts (radical, cationic, anionic) at homopolymerization were unsuccessful. We therefore turned our attention to the preparation of a phosphazene bearing a styrene, rather than an  $\alpha$ -methylstyrene, substituent. Since anionic polymerization of styrene is initiated by organolithium reagents, indirect routes were explored. Aryl pentafluorocyclotriphosphazenes with carbonyl groups in the para position (3,4) were prepared from para-bromobenzaldehyde and para-bromoacetophenone by first converting the carbonyl to the acetal and then preparing the lithium



reagent. After coupling with  $\text{N}_3\text{P}_3\text{F}_6$  the protecting group was removed to give 3 and 4. However, the Wittig reaction on 3 gave only degradation products as did the reduction of 4 to the alcohol. An alternative route was developed starting from an  $\alpha$ -methoxystyryl phosphazene (5). Interestingly, 5 was unreactive with respect to borohydride or aluminum hydride addition. Fortunately, hydrostanation of 5 proceeded smoothly and  $\beta$ -elimination of tributyltinmethoxide from the resulting organotin derivative occurred spontaneously on silica gel giving rise to the styrylpentafluorocyclotriphosphazene (6).



The  $^{13}\text{C}$  chemical shift of the vinylidene carbon in 6 occurs further downfield than even 4-nitrostyrene, making 6 amongst the most electron deficient styrene derivatives known. Polymerization of 6 via radical initiation proceeded smoothly to give the homopolymer (7), a polymer having a carbon chain backbone and an inorganic surface. The molecular weights of 7 are 102,000 and



374,000 for  $M_n$  and  $M_w$  respectively. Thermal decomposition of 7, as monitored by TGA, starts about  $80^\circ$  higher than that for polystyrene and the polymer retains fifteen to thirty percent of its original weight at temperatures of  $900^\circ$  to  $1100^\circ\text{C}$ . This observation suggests the possibility of thermal decomposition of carbon chain polymers with inorganic surfaces as a potentially new route to ceramic materials.

#### Acknowledgements

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#### References

- 1) Allen, C.W. J. Polym. Sci., Polym. Symp. 1983, 70, 99.
- 2) Allcock, H.R.; Hymer, W.C.; Austin, P.E. Macromolecules 1983, 16, 1401.
- 3) Kumar, D.; Fohlen, G.M.; Parker, J.A. J. Polym. Sci., Polym. Chem. Ed. 1984, 22, 1411.
- 4) Ramachandran, K.; Allen, C.W. Inorg. Chem. 1983, 22, 1445.
- 5) Allen, C.W.; Bright, R.P. Inorg. Chem. 1983, 22, 1291.
- 6) Allen, C.W.; Desorcie, J.L.; Ramachandran, K. J. Chem. Soc., Dalton Trans. 1984, 2843.
- 7) Dupont, J.G.; Allen, C.W. Macromolecules 1979, 12, 169.
- 8) Dupont, J.G.; Allen, C.W. Inorg. Chem. 1978, 17, 3093.
- 9) Allen, C.W.; Bright, R.P. Inorg. Chim. Acta 1985, 99, 107.



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